

Name \_\_\_\_\_

AP Chemistry

## Inorganic Chemistry Nomenclature Practice

### Part I

Name each of the following compounds.

- |  |  |
|--|--|
| 1) NaBr—Sodium bromide   | 9) N <sub>2</sub> O <sub>4</sub> —Dinitrogen tetroxide                   |
| 2) Rb <sub>2</sub> O—Rubidium oxide                                    | 10) ICl <sub>3</sub> —Iodine trichloride                                 |
| 3) CaS—Calcium sulfide   | 11) SO <sub>2</sub> —Sulfur dioxide                                      |
| 4) AlI <sub>3</sub> —Aluminum iodide                                   | 12) P <sub>2</sub> S <sub>3</sub> —Diphosphorus trisulfide               |
| 5) BaSO <sub>3</sub> —Barium sulfite                                   | 13) HC <sub>2</sub> H <sub>3</sub> O <sub>2</sub> —Acetic acid           |
| 6) NaNO <sub>2</sub> —Sodium nitrite                                   | 14) H <sub>2</sub> SO <sub>4</sub> —Sulfuric acid                        |
| 7) KMnO <sub>4</sub> —Potassium permanganate                           | 15) HClO—Hypochlorous acid   |
| 8) K <sub>2</sub> Cr <sub>2</sub> O <sub>7</sub> —Potassium dichromate | 16) Pb <sub>3</sub> (PO <sub>4</sub> ) <sub>2</sub> —Lead (II) phosphate |

### Part II

Write the formula for the following compound.

- 1) Zinc chloride—ZnCl<sub>2</sub>
- 2) Tin (IV) fluoride—SnF<sub>4</sub>
- 3) Calcium nitride—Ca<sub>3</sub>N<sub>2</sub>
- 4) Aluminum sulfide—Al<sub>3</sub>S<sub>2</sub>
- 5) Mercury (I) selenide—Hg<sub>2</sub>Se

- 6) Silver iodide— $\text{AgI}$
- 7) Ammonium hydrogen phosphate— $(\text{NH}_4)_2\text{HPO}_4$
- 8) Mercury (I) sulfide— $\text{Hg}_2\text{S}$
- 9) Silicon dioxide— $\text{SiO}_2$
- 10) Sodium sulfite— $\text{Na}_2\text{SO}_3$
- 11) Aluminum hydrogen sulfate— $\text{NH}_4\text{HSO}_4$
- 12) Nitrogen trichloride— $\text{NCl}_3$
- 13) Hydrobromic acid— $\text{HBr}$
- 14) Bromous acid— $\text{HBrO}_2$
- 15) Perbromic acid— $\text{HBrO}_4$
- 16) Potassium hydrogen sulfide— $\text{KHS}$
- 17) Calcium iodide— $\text{CaI}_2$
- 18) Cesium perchlorate— $\text{CsClO}_4$